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Those are just the orbitals with the atoms by themselves. Now these orbitals in the middle have to do with interference. They have to do with how the atomic orbitals coming together. These are called molecular orbitals. So what molecular orbitals do is they predict where electrons are going to be

in the entire molecule, not just for the atom.

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Hybridization is essential for understanding the geometry of covalent bonds. Here a covalent bond is created by the overlapping of two atomic orbitals that are not fully occupied. Hybrid

atomic orbitals mix together atomic orbitals to form an equal number of new hybrid atomic orbitals with a varied shape. For example, a 2s and 2p orbital would combine to form sp hybrids of a varied molecular ...

Molecular Orbitals and Hybridizations - Organic Chemistry

...

The manner in which atomic orbitals overlap to form molecular orbitals is actually more complex than the localized examples given above. These are useful models for explaining the structure and reactivity of many organic compounds, but modern molecular orbital theory involves the creation of an orbital correlation diagram.

Molecular Structure & Bonding

History. In 1952, Kenichi Fukui published a paper in the Journal of Chemical Physics titled "A molecular theory of reactivity in aromatic hydrocarbons." Though widely criticized at the time, he later shared the Nobel Prize in Chemistry with Roald Hoffmann for his work on reaction mechanisms. Hoffman's work focused on creating a set of four pericyclic reactions in organic chemistry, based on orbital symmetry, which he coauthored with Robert Burns Woodward, entitled "The Conservation of ...

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For every molecule, two of its molecular orbitals will be of particular interest: 1) the highest energy orbital from among the occupied orbitals known as the HOMO and 2) the lowest energy orbital from among the unoccupied orbitals

called the LUMO. 1.09 Hybridized Atomic Orbitals On the left is the electron configuration of carbon.

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graduate level course in physical organic chemistry I would use this as one of the suggested titles. It is a must for the

practicing physical organic chemist. It is sufficiently complex to give an understandability to very difficult subject material.