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The reaction is first order with respect to each of the reactants, and at 298 K, the rate constant is $1.02 \times 10^{-6} \text{ L mol}^{-1} \text{ s}^{-1}$. Calculate the initial rate of the reaction in a 4.0 L reaction vessel containing 0.60 moles of iso-butane and 0.00010 moles of bromine atoms. A) $1.5 \text{ mol L}^{-1} \text{ s}^{-1}$: B) $3.8 \text{ mol L}^{-1} \text{ s}^{-1}$: C) $2.5 \times 10^{-11} \text{ mol L}^{-1} \text{ s}^{-1}$: D) $2.5 \times \dots$