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Using the second law of thermodynamics, explain why heat flows from a hot body to a cold body but not from a cold body to a hot body. One test of the spontaneity of a reaction is whether the entropy of the universe increases: $\Delta S_{\text{univ}} > 0$. Using an entropic argument, show that the following reaction is spontaneous at 25°C:

19.E: Chemical Thermodynamics (Exercises) - Chemistry ...

chapter. Lecture Outline 19.1 Spontaneous Processes1,2,3,4 • Chemical thermodynamics is concerned with energy relationships in chemical reactions. • We consider enthalpy. • We also consider entropy in the reaction. • Recall the first law of thermodynamics: energy is conserved. $\Delta E = q + w$

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Chemical Thermodynamics ΔS For Phase Change Consider the melting of ice • At 1 atm and 0 °C ice and water are in equilibrium with each other. • If we melt one mole of ice at 0 °C/1 atm to form 1 mole of water still at 0 °C • We need to give the water heat that would be equal to ΔH_{fusion} . • If the process is done extremely slowly the process

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Chapter 19 notes - University of Massachusetts Boston

AP Chemistry Chapter 19 Chemical Thermodynamics - 1 - Chapter 19. Chemical Thermodynamics . 19.2 Entropy and the Second Law of Thermodynamics . Entropy Change • Entropy, S, is a thermodynamic term that reflects the disorder, or randomness, of the system. • The more disordered, or random, the system is, the larger the value of S.

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Chapter 19 Chemical Thermodynamics Chemical Thermodynamics Enthalpy • A thermodynamic

quantity that equal to the internal energy of a system plus the product of its volume and pressure exerted on it by its surroundings; "Enthalpy is the amount of energy in a system capable of doing mechanical work" Using the symbol H for the enthalpy: $H=E+pV$ Chemical Thermodynamics

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