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Ideal Gas Law and Stoichiometry Use the following reaction to answer the next few questions:  $2 \text{ C}_8\text{H}_{18}(l) + 25 \text{ O}_2(g) \rightarrow 16 \text{ CO}_2(g) + 18 \text{ H}_2\text{O}(g)$  The above reaction is the reaction between gasoline (octane) and oxygen that occurs inside automobile engines.

**Gas Laws STUDY GUIDE Due: February 12th**

The three fundamental gas laws discover the relationship of pressure, temperature, volume and amount of gas. Boyle's Law tells us that the volume of gas increases as the pressure decreases. Charles' Law tells us that the volume of gas increases as the temperature increases. And Avogadro's Law tell us that the volume of gas increases as the amount of gas increases. The ideal gas law is the combination of the three simple gas laws.

**Gas Laws: Overview - Chemistry LibreTexts**

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Practice Test: Gas Laws. 11. Zinc metal is added to hydrochloric acid to generate hydrogen gas and is collected over a liquid whose vapor pressure is the same as pure water at  $20.0^\circ\text{C}$  (18 torr). The volume of the mixture is 1.7 L, and its total pressure is 0.810 atm.

**Practice Test: Gas Laws**

The formula of this law is as follows:  $(8.4.6) P V = n R T$ . In this equation, P is pressure, V is volume, n is amount of moles, and T is temperature. R is called the ideal gas law constant and is a proportionality constant that relates the values of pressure, volume, amount, and temperature of a gas sample.

**8.4: Gas Laws - Chemistry LibreTexts**

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