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Mole Ratio Practice Problems Chapter 12 Stoich Limiting Reactant How To: Find Limiting Reagent (Easy steps w/practice problem) How to round to significant figures (s.f.)

Stoichiometry Made Easy: The Magic Number Method

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Answer Key Chapter 12: Stoichiometry Mole Ratios Questions 1. Aluminum reacts with oxygen to produce aluminum oxide as follows: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ a. If you use 2.3 moles of Al, how many moles of Al_2O_3 can you make? b. If you want 3.9 moles of Al_2O_3 , how many moles of O_2 are needed? 2.

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Merely said, the chapter 9 stoichiometry work answers is universally compatible past any devices to read. chapter 9 stoichiometry work answers CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: $\text{C}_3\text{H}_4(\text{g}) + \text{xO}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ 4 a.

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