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General form of an equilibrium reaction: $aA + bB \rightleftharpoons cC + dD$ "A & B" are reactants, "C & D" are products, and "a, b, c, d" are coefficients. To determine the equilibrium constant (K) from the general model: $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ If a K-value greater than 1 is obtained, more products than reactants at equilibrium. If a K-value less than one is obtained, there are more reactants than products at equilibrium.

Chemical Equilibrium Quiz - Softschools.com

The following chemical reaction has reached equilibrium. $2NO(g) + O_2(g) \rightleftharpoons 2NO_2(g)$ Calculate the equilibrium constant, K_{eq} , if the concentrations of reactants and products are $1.26 \times 10^{-10} M$ for NO, $2.28 \times 10^{-3} M$ for O_2 , and $9.32 \times 10^{-6} M$ for NO_2 .

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Test prep MCAT Chemical processes Equilibrium. Equilibrium. Practice: Equilibrium questions. This is the currently selected item. Reactions in equilibrium. Le Chatelier's principle. Changes in free energy and the reaction quotient. Standard change in free energy and the equilibrium constant.

Equilibrium questions (practice) | Khan Academy

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Equilibrium Name _____ MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) At equilibrium, _____. A) the rates of the forward and reverse reactions are equal B) the rate constants of the forward and reverse reactions are equal

A.P. Chemistry Practice Test - Ch. 13: Equilibrium ...

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Equilibrium Class 11 MCQs Questions with Answers. Question 1. The solubility of $\text{Ca}_3(\text{PO}_4)_2$ in water is y mol/L. Its solubility product is: (a) $6y^2$ (b) $36y^4$ (c) $64y^5$ (d) $108y^5$. Answer: (d) $108y^5$ Explanation: $= (3y)^3 \times (2y)^2 = 27y^3 \times 4y^2 = 108y^5$

MCQ Questions for Class 11 Chemistry Chapter 7 Equilibrium ...

Unformatted text preview: Chemical Equilibrium Use the website linked below to answer the following questions, submit when finished 1. At chemical equilibrium, what is the relationship between the rate of the forward reaction and the rate of the reverse reaction? A reversible reaction can proceed in both the forward and backward directions.

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Chemistry Quiz Questions Chemical Equilibrium - Chemistry Questions 31-60 . 31) The steady state from which no spontaneous change takes place is called what? Answer: Equilibrium. 32) What is the color of ammonium chloride? Answer: White. 33) The chemical reactions which never go to completion are called what? Answer: Reversible reactions.

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NCERT Solutions for Class 11 Chemistry Chapter 7 Short Answer Type Questions

Question 1. The following concentration were obtained for the formation of NH_3 from N_2 and H_2 at equilibrium at 500 K. $[\text{N}_2(\text{g})] = 1.5 \times 10^{-2} \text{ M}$ $[\text{H}_2(\text{g})] = 3.0 \times 10^{-2} \text{ M}$ $[\text{NH}_3] = 1.2 \times 10^{-2} \text{ M}$. Calculate equilibrium constant.

**NCERT Solutions for Class 11
Chemistry Chapter 7 Equilibrium**

Preview this quiz on Quizizz. A chemist is working with a liquid substance that tastes bitter, feels slippery, and conducts an electric current. Which of the following is the chemist most likely

working with?

**AP Chemistry--Acid Base Equilibria
Quiz - Quizizz**

A reversible chemical process is considered in equilibrium when the rate of the forward reaction equals the rate of the reverse reaction. The ratio of these reaction rates is called the equilibrium constant. Test your knowledge about equilibrium constants and their use with this ten question equilibrium constant practice test. Answers appear at the end of the test.