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A buffer is an aqueous mixture of a weak acid and its conjugate base, or a weak base and its conjugate acid. A buffer has a very stable pH. When small amounts of an acid or a base are added to a buffer solution, the pH of the solution changes very little. In many chemical and biochemical systems, buffers are critical.

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A buffer is a mixture of a weak acid and its conjugate base, or a weak base and its conjugate acid. A buffer's function is to absorb acids (H + or H 3 O + ions) or bases (OH - ions) so that the pH of the system changes very, very little. In many systems, buffers are critical.

Buffers - Vernier

A buffer is the combination of a weak acid and a salt of the weak acid. Acetic acid and sodium

acetate are an example of this kind of buffer pair. Buffers resist changes in pH upon the addition of small amounts of H + or OH - ions. The dissociation equation for acetic acid contains both of the buffer components, HC 2 H 3 O 2 and C 2 H 3 O 2-:

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