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 principle can be
 applied to understand
 the effect of change in

pressure on the
 systems at equilibrium
 as follows. 1) When the
 partial pressure of any
 of the gaseous
 reactants or of the
 products is increased ,
 the position of
 equilibrium is shifted
 so as to decrease its
 partial pressure .

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 all reacting species
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states that a system at
equilibrium will
respond to a stress on
the system in such a
way so as to relieve
the stress and
establish a new
equilibrium. The
system will have one
reaction

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state Le Chatelier's
principle when a

physical or chemical
system at equilibrium
is disturbed by
applying a stress
(changing the
temperature, pressure
or concentration), it
acquires a new
equilibrium so as to
relieve the stress.

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The Le Chatelier's
principle has very
important and wide
application in chemical
reactions that are both
exothermic and
endothermic in nature.
It aims at maintaining
an equilibrium position
for all reacting species
even one any factor is
altered so that the
reaction can proceed
to completion.

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has very important and
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even one any

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has been observed
that when a reaction at
equilibrium is disturbed
by changing either the
concentration of one of
the chemical
components, the total
pressure, or the
temperature, the
reaction will respond
by shifting its
equilibrium position so
as to counteract the
effect of the
disturbance.

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Le Chatelier's Principle
states that if a stress is
applied to a reversible
reaction at equilibrium,
the reaction will
undergo a shift in order
to re-establish its
equilibrium. Consider
the following
exothermic reversible
reaction at equilibrium:
 $2A \rightleftharpoons B + C$

**12: Equilibrium and
Le Chatelier's
Principle
(Experiment ...**

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Le Chatelier's principle can be used to predict the behavior of a system due to changes in pressure, temperature, or concentration. Le Chatelier's principle implies that the addition of heat to a reaction will favor the endothermic direction of a reaction as this reduces the amount of heat produced in the system.