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Boyle's Law Worksheet Gas law packet answers 1. Boyles' Law Use Boyles' Law to answer the following questions: 1) 1.00 L of a gas at standard temperature and

pressure is compressed to 473 mL. Gas law packet answers - SlideShare Gas Laws Worksheet $\text{atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760 .0 \text{ torr}$ Boyle's Law Problems: 1.

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Boyles Law Packet Answers Boyle's Law states that the volume of a gas varies inversely with its pressure if temperature is held Solve the following problems (assuming constant temperature). Show your work on the back of the worksheet and place your answer. Go to Boyle's Law Answers: #1 - 15 · Go to Boyle's Law Answers: #16 - 31.

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law combined gas law

$p_1 v_1 = p_2 v_2$

the pressure of a gas is

directly proportional to

the kelvin temperature

if the volume is kept

constant. the volume

of a

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Solution for Boyle's law

states a volume of a

gas varies inversely

proportional to its

pressure at a constant

temperature. True or

false? Give an

example?

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 questions: 1) 1.00 L of a
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 temperature and
 pressure is
 compressed to 473 mL.
 What is the new
 pressure of the gas?
 2.11 atm 2) In a
 thermonuclear device,
 the pressure of 0.050
 liters of gas within the
 bomb casing reaches
 4.0×10^6 atm.

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68kit.dmitrichavkerovnews.me We know that
 pressure and volume
 are inversely related;
 as one decreases, the
 other increases.
 Pressure is decreasing
 (from 2.44 atm to 1.93
 atm), so volume should
 be increasing to
 compensate, and it is
 (from 4.01 L to 5.07 L).
 So the answer makes
 sense based on Boyle's
 law.

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A gas law is a simple
 mathematical formula
 that allows you to
 model, or predict, the
 behavior of a gas. This
 particular gas law is
 called Boyle's law,
 after the English
 scientist Robert Boyle,
 who first announced it

in 1662. Figure
\\(\\PageIndex{1}\\)
shows two
representations of how

Boyle's law works.
Figure
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