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2) Solve for the heat required to change the water into steam (no change in temp). 3) Calculate the heat required to change the temperature of the steam from 100.0 oC to 110.0 oC. 4) To get the heat required for the whole process, ____ the calculated heats from above. Q = m x ΔH vap Q = 1000. g x 2260 J/g = 2,260,000 J Q = m x C x Δt

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