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454 CHAPTER 15 Formula Acid name HF hydrofluoric acid HCl hydrochloric acid HBr hydrobromic acid HI hydriodic acid H<sub>2</sub>S hydrosulfuric acid TABLE 15-1 Names of Binary Acids FIGURE 15-2 A strip of pH paper dipped into vinegar turns red, showing that vinegar is an acid.

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Chapter Fifteen: Acid-Base Titration and pH If you look around the room while titrations are happening, there are people standing like this: and the best part is that you probably will too, because it works.

### Chapter Fifteen [Acid-Base Titration and pH]

This chapter 15 test acids and bases answers, as one of the most working sellers here will categorically be in the middle of the best options to review.

### Chapter 15 Test Acids And Bases Answers

Chapter 14/15 Equilibrium and Acid-Base Practice Test. Multiple Choice (Choose the best answer.-- Answers Found At End of Test) \_\_\_\_ 1. The reaction: 3 CuCl<sub>2</sub> (aq) + 2 Al (s) 3 Cu (s) + 2 AlCl<sub>3</sub> (aq) has the following K<sub>c</sub> expression:. a) [CuCl<sub>2</sub>]<sup>3</sup> [Al]<sup>2</sup> / { [Cu]<sup>3</sup> [AlCl<sub>3</sub>]<sup>2</sup> } b) [Cu]<sup>3</sup> [AlCl<sub>3</sub>]<sup>2</sup> / { [CuCl<sub>2</sub>]<sup>3</sup> [Al]<sup>2</sup> } c) [CuCl<sub>2</sub>]<sup>3</sup> / [Cu]<sup>3</sup> d) [Al]<sup>2</sup> / [AlCl<sub>3</sub>]<sup>2</sup> e) None of these

### Chapter 14/15 Test--Equilibrium and Acids/Bases

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1. Aqueous solutions of acids have a sour taste. 2. Acids change the color of acid-base indicators. 3. Some acids react w/active metals to release hydrogen gas, H<sub>2</sub>. 4. Acids react with bases to produce salts and water. (salt- the reaction products are water and an ionic compound called a salt). 5. Some acids conduct electric current. 6 ...

#### **Modern Chemistry Chapter 15 Acids and Bases | StudyHippo.com**

Answers to Chapter 15 & 16 Study Questions. 1. strong acid + strong base: steepest at equivalence, low pH at start depending on concentration of acid, equivalence at pH 7, no buffering. weak acid + strong base: less steep at equivalence than strong acid, low pH at start depending on and

concentration of acid, equivalence at basic pH, buffering. weak base + strong acid: less steep at equivalence than strong base, high pH at start depending on and concentration of base, equivalence at acid pH, ...

#### **Answers to Chapter 15 & 16 Study Questions - Chemistry ...**

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