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Section 92 Classifying Chemical Reactions

a chemical reaction that occurs when a substance reacts with oxygen, releasing energy in the form of heat and light decomposition reaction a chemical reaction that occurs when a single compound breaks down into two or more elements or new compounds

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The term oxidation was originally used to describe chemical reactions involving O 2, but its meaning has evolved to refer to a broad and important reaction class known as oxidation-reduction (redox) reactions. A few examples of such reactions will be used to develop a clear picture of this classification.

4.2: Classifying Chemical Reactions - Chemistry LibreTexts

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Decomposition Single & Double Replacement Combustion Precipitation. Terms in this set (24) Synthesis Reaction. a reaction in which two or more substances reacts to form a single substance.

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Most chemical reactions can be classified into one or more of five basic types: acid-base reactions, exchange reactions, condensation reactions (and the reverse, cleavage reactions), and oxidation-reduction reactions. The general forms of these five kinds of reactions are summarized in Table 7.10.1, along with examples of each.

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By the end of this section, you will be able to: Define three common types of chemical reactions (precipitation, acid-base, and oxidation-reduction) Classify chemical reactions as one of these three types given appropriate descriptions or chemical equations; Identify common acids and bases

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Most chemical reactions can be classified into one or more of five basic types: acid-base reactions A reaction of the general form acid + base → salt.,

exchange reactions A chemical reaction that has the general form $AB + C \rightarrow AC + B$ or $AB + CD \rightarrow AD + CB$., condensation reactions A chemical reaction that has the general form $A + B \rightarrow AB$. Condensation reactions are the reverse of cleavage reactions.

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Section 9.2 Classifying Chemical Reactions: • combustion reaction (p. 290) • decomposition reaction (p. 292) • double-replacement reaction (p. 296) • precipitate (p. 296) • single-replacement reaction (p. 293) • synthesis reaction (p. 289) Section 9.3 Reactions in Aqueous Solutions:

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