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$\text{Zn(s)} + \text{Cu}^{2+} + 4 \text{Zn}^{*+} + \text{Cu(s)}$ An electrolytic cell based on the reaction repre- sented above was constructed from zinc and copper half-cells. The observed voltage was found to be 1.00 volt instead of the standard cell potential, E° , of 1.10 volts. Which of the following could correctly account for this observation?

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Advanced Placement Chemistry: 1994 Free Response Questions. 1) $\text{MgF}_2(\text{s}) \rightleftharpoons \text{Mg}^{2+}(\text{aq}) + 2 \text{F}^{-}(\text{aq})$ In a saturated solution of MgF_2 at 18 C, the concentration of Mg^{2+} is 1.21×10^{-3} molar. The equilibrium is represented by the equation above.
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