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Chapter 13 Study Guide Gases

Chapter 13 – Gases 195 Exercise 13.3 – Equation Stoichiometry: Iron is combined with carbon in a series of reactions to form pig iron, which is about 4.3% carbon. $2C + O_2 \rightarrow 2CO$ $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$ $2CO + C \rightarrow 3C$ (in iron) CO_2 Pig iron is easier to shape than pure iron, and the presence of carbon lowers its melting point

Chapter 13 - Gases

Chapter 13 Study Guide: Gases. Students should be able to... · Explain the kinetic theory of matter as it applies to gases. · Identify the properties of an ideal gas vs. a real gas. · Know units of pressure in atm, torr, mm Hg, kPa, psi, and in Hg. · Calculate using Boyle's Law, Charles' Law, Gay-Lussac's Law, Combined Gas Law, and Ideal Gas Law (using 0.0821 for R), and Graham's Law.

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194 Study Guide for An Introduction to Chemistry Section Goals and Introductions Section 13.1 Gases and Their Properties Goals • To describe the particle nature of both real and ideal gases. • To describe the properties of gases that can be used to explain their characteristics: volume, number of particles, temperature, and pressure.

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A gas can flow into a space occupied by another gas. 9. The diffusion of a gas is caused by the random motion of the particles of the gas. 10. Lighter gas particles diffuse less rapidly than do heavier gas particles. 11. During effusion, a gas escapes through a tiny opening into a vacuum. 12.

13 STUDY GUIDE FOR CONTENT MASTERY 13 STUDY GUIDE FOR ...

Chapter 13 Study Guide Answers 1. Describe the assumptions/postulates of the kinetic-molecular theory of gases: a. Gases are composed of tiny particles in constant rapid, ("random" or "straight-lined"?) motion. b. Gases are separated by relatively huge distances. The volume of the particles is essentially zero. c.

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Chapter 13 - Gases 190 Study Guide for An Introduction to Chemistry Section Goals and Introductions Section 131 Gases and Their Properties Goals To describe the particle nature of both real and ideal gases To describe the properties of gases that can be used to explain