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Collision Theory Gizmo : Lesson Info : ExploreLearning

The Collision Theory Gizmo™ allows you to experiment with several factors that affect the rate at which reactants are transformed into products in a chemical reaction. You will need blue, green, and orange markers or colored pencils for the first part of this activity. 1. Look at the key at the bottom of the SIMULATION pane.

Collision Theory Notes - Name Date Student Exploration ...

Collision theory states that there needs to be enough energy to break the bonds between molecules and smash the electron clouds together Why does increasing temperature speed up a chemical reaction? The energy of the collisions increase because the temperature raises the average kinetic energy of the particles.

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Collision theory. all reactions require activation energy, so reactions will only occur when atoms, and compounds collide with enough energy. How does increased concentration increase the rate of reaction? More molecules are present for collisions to occur. With higher concentration, more reactants, on average, will have enough energy to proceed forward.

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