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in Schroeder), $\Delta H_{H_2O} = 285.83 \text{ kJ mol}^{-1}$ and $\Delta H_{H_2O} = 890.36 \text{ kJ mol}^{-1}$ (12) This time, the final volume is $\frac{1}{3}$ of the initial volume, since the 2 moles of water has condensed out as liquid with negligible volume compared to the gases. Thus the atmosphere does work $P\Delta V = RT\Delta n$ (13) $= 8.31 \text{ J K}^{-1} (298 \text{ K}) (2) = 4.953 \text{ kJ}$ (15)

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