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According to modern atomic theory, we cannot be sure of the exact location of electrons in an atom. We predict that electrons will be relatively close to the nucleus (because the electrons are negatively charged and the nucleus is positively charged).

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$N=3$, $L=1$, $M=1$, $M=1/2$. This electron is found within the 2p orbital, which has the dumbbell shape, primarily based on the z axis. This electron would share this orbital with another electron, but would have a negative spin.

H-Atom Analogy Lab Discussion Questions | Atomic Orbital ...

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(d) If we count the number of dots in each spherical shell, we obtain the total probability of finding the electron at a given value of r . Because the surface area of each shell increases more rapidly with increasing r than the electron probability density decreases, a plot of electron probability versus r (the radial probability) shows a peak. This peak corresponds to the most probable radius for the electron, 52.9 pm, which is exactly the radius predicted by Bohr’s model of the hydrogen ...

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