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214 Chapter 8 Ionic Compounds Figure 8-3 In the formation of a negative ion, a neutral atom gains one or more electrons. Again, note that in the neutral atom the number of protons equals the

number of electrons. However, the ion contains more electrons than protons, making this over-all charge on this ion negative. e

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Answer Key 8.2 Ionic Formulas

Compounds are Neutral They have no net charge so you must have enough cations and anions to equal zero. Na⁺ and Cl⁻ make NaCl since +1 and -1 = 0 + Na and CN⁻ make NaCN since +1 and -1 = 0 2+ Ba and Cl⁻ make BaCl₂

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8. In a crystal lattice of an ionic compound, a. ions of a given charge are clustered together, far from ions of the opposite charge. ons are surrounded by ions of the opposite charge. c. a sea of

electrons surrounds the ions. d. neutral molecules are present. Chemistry: Matter and Change Chapter 8 Study Guide for Content Mastery 44

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Chapter 8 - Ionic Compounds - ProProfs Quiz

Date Class CHAPTER 8 CHAPTER ASSESSMENT Applying Scientific Methods A college chemistry student is studying the properties of four unknown compounds, W,x,y and Z. She has been informed that one of them is ionic and that the other three are covalent Of the latter, she has been told that the attractions between the formula units are dispersion forces in one case, hydrogen bonds in another, and covalent

(network) bonds in another.

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Section 8.2 What is an ionic bond? In your textbook, read about forming ionic bonds and the characteristics of ionic compounds. Circle the letter of the choice that best completes the statement or answers the question. 1. An ionic bond is a. attraction of an atom for its electrons. b. attraction of atoms for electrons they share.

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