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A covalent bond is formed when two atoms share a pair of electrons. Covalent bonding occurs in most non-metal elements, and in compounds formed between non-metals. These shared electrons are found...

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Covalent bonds form when electrons are shared between atoms and are attracted by the nuclei of both atoms. In pure covalent bonds, the electrons are shared equally. In polar covalent bonds, the electrons are shared unequally, as one atom exerts a stronger force of attraction on the electrons than the other.

Covalent Bonding | General Chemistry - Lecture & Lab

Covalent bonds are commonly found in carbon-based organic molecules, such as DNA and proteins. Covalent bonds are also found in inorganic molecules such as H₂O, CO₂, and O₂. One, two, or three pairs of electrons may be shared between two atoms, making single, double, and triple bonds, respectively.

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A covalent bond, also called a molecular bond [citation needed], is a chemical bond that involves the sharing of electron pairs between atoms.These electron pairs are known as shared pairs or bonding pairs, and the stable balance of attractive and repulsive forces between atoms, when they share

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