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Spectroscopy received. So, if IE is the ionization energy of the electron and KE is the kinetic energy with which it leaves the atom, we have $E_{\text{photon}} = \text{IE} + \text{KE}$ or, upon rearranging the equation, $\text{IE} = E_{\text{photon}} - \text{KE}$. Thus, we can find the ionization energy, IE, if we

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Figure 2: A simulated photoelectron spectrum of the hypothetical atom in Figure 1.
Critical Thinking Questions 3. Use the data presented in Figure 1 to verify that the IE of the ejected electron is 28.6 MJ/mole.

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Page 7/29. Download File PDF Chemactivity 8 Photoelectron Answers of Ne because both atoms have eight electrons in the 2nd shell and F⁻ has a core charge of +7 whereas Ne has a core charge of +8. 6.IE of Kr > IE of Br

because they are in the same valence shell and Kr has the higher core charge (+8 vs. +7).

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ChemActivity 8 Photoelectron Spectroscopy 47Figure 2: A simulated photoelectron spectrum of the hypothetical atom in Figure 1.Critical Thinking Questions 3. Use the data presented in Figure 1 to verify that the IE of the ejected electron is 28.6 MJ/mole. 4.

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