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Answers: Moles and Stoichiometry Practice Problems 1) How many moles of sodium atoms correspond to 1.56x10²¹ atoms of sodium? 1.56 -x 10²¹ atoms Na x 1 mol Na = 2.59 x 10³ mol Na 236.022 x 10 atoms Na 2) Determine the mass in grams of each of the following: a. 1.35 mol of Fe 1.35 mol Fe x 55.845 g Fe = 75.4 g Fe 1 mol Fe b. 24.5 mol O

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Answers to Stoichiometry Practice Free Response Test: 1. 30.40 g. N X 1 mol N/14.01g N = 2.17 moles N. 69.60 g O X 1 mol. O /16.00 g O = 4.35 moles O. 4.35/2.17. = 2.00 2.17/2.17 = 1.00 So O is 2X more present than N. Empirical formula must be NO. 2. Mass.

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Chapter 12 Stoichiometry Test Answer Key

Practice Test Ch 3 Stoichiometry Name Per. Remember it is a MC test, use the answers ... Practice Test Ch3 Stoichiometry (page 3 of 3) 1. d It might be easiest to balance the equation with mostly whole numbers: 2 NH₃ + 7/2 O₂ → 2NO₂ + 3H₂O ... 7. c First you must realize this is a limiting reactant problem. Found: 3 Jan 2020 | Rating: 89/100

Stoichiometry Practice Problems Level 1 Answers

20 Then do some stoichiometry using “easy math” 16 g of methane (MM = 16) is 1 mole and 1 mole of methane will produce 1 mole of CO₂ = 44 g, and 2 moles of H₂O which is 36 g for a total of 80 g 4. d Balance: C₃H₈ + 5O₂ → 3CO₂ + 4H₂O 5. d Balance: 2KClO₃ → 2KCl + 3O₂

Practice Test Ch 3 Stoichiometry Name Per

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Chemical Stoichiometry Test Answers

Stoichiometry Practice Test Short Answer: Aluminum bromide can be prepared by the reaction of aluminum metal with bromine gas shown by the equation: 2 Al + 3 Br₂ → 2 AlBr₃ Now suppose that 5.6 mol of aluminum reacts with 4.4 mol of bromine. 1. Calculate the mass of aluminum bromide that can be produced from 5.6 mol of Al. 2.

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