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Answer: (a) SrSO₄(s) ⇌ Sr²⁺(aq) + SO₄²⁻(aq) At equilibrium: [Sr²⁺] = X M = [SO₄²⁻] X₂ = K_{sp} = 7.6 × 10⁻⁷ X = 8.7 × 10⁻⁴ mol/L, solubility of SrSO₄ (b) SrF₂(s) ⇌ Sr²⁺(aq) + 2 F⁻(aq) At equilibrium: [Sr²⁺] = X M = [F⁻] = 2X M K_{sp} = [Sr²⁺][F⁻]² = (X)(2X)² = 7.9 × 10⁻¹⁰ X = 5.8 × 10⁻⁴ mol/L, solubility of SrF₂

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NaOH(aq). (i) Calculate the number of moles of NaOH(aq) added.

molNaOH= 6.55 mL × 1 L 0.435 mol NaOH 1,000 mL 1 L × molNaOH= 2.85 × 10⁻³mol NaOH. One point is earned for the correct number of moles of NaOH. AP®CHEMISTRY 2005 SCORING GUIDELINES (Form B) Copyright © 2005 by College Board. All rights reserved.

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By stoichiometry, [I⁻] = 2 × [Pb²⁺] , thus [I⁻] = 2 × (1.3 × 10⁻³) = 2.6 × 10⁻³M. One point is earned for the correct concentration. (iii) Calculate the value of the equilibrium constant, K. sp. 2K. sp= [Pb. 2+][I⁻] = (1.3 × 10⁻³)(2.6 × 10⁻³)². = 8.8 × 10⁻⁹. One point is earned for a value of K.

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Answer: (a) S = S (dia.) - S (graph.) = (0.58 - 1.37) cal/K = -0.79 cal/K CO₂ C(dia.) + O₂ H = + 94.03 kcal/mol C(graph.) + O₂ CO₂ H = - 94.48 kcal/mol C(graph.) C(dia.) H = -0.45 kcal/mol (b) G = H - T S = -450 - (298)(-0.79) cal = -223.52 cal/mol; a G < 0 indicates feasible conditions

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Copyright © 2005 by College Entrance Examination Board. All rights reserved. Available at apcentral.collegeboard.com. 1 HC₃H₅O₂(aq) ⇌ C₃H₅O₂⁻(aq) + H⁺(aq) K_a= 1.34 × 10⁻⁵. Propanoic acid, HC₃H₅O₂, ionizes in water according to the equation below. a) write the equilibrium-constant expression for the reaction.

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equation or an answer obtained in the previous part(s) of the question. • Make use of the Nernst equation to calculate the E for a galvanic cell. • Encourage students to fully establish equations before making substitutions. ... 2004 AP Chemistry Free Response Questions

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