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solution and diluting it to 100 mL. What absorbance do you expect of the final diluted solution, if you made it up correctly, at $\lambda = 259 \text{ nm}$? Dilution gives $C_1 \times V_1 = C_2 \times V_2$; $0.025 \text{ mL} \times 0.1 \text{ M} = 100 \text{ mL} \times C_{\text{ATP}}$ $C_{\text{ATP}} = 0.025 \text{ mL} \times 0.1 \text{ M} / 100 \text{ mL} = 25 \times 10^{-6} \text{ M}$ $A = \epsilon C l = 15.4 \times 10^3 \text{ M}^{-1} \times 25 \times 10^{-6} \text{ M} = 0.385$ 3. A.