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Answers reactions, condensation reactions (and the reverse, cleavage reactions), and oxidation-reduction reactions. The general forms of these five kinds of reactions are summarized in Table, along with examples of each.

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STUDY GUIDE For each of the following chemical reactions, write a word equation, a skeleton equation, and a balanced chemical equation. Be sure to show the state of each reactant and product. If you need more help writing formulas or determining the state of a substance, refer to Chapters 7 and 8 and

the periodic table on pages 178—179. 17.

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Know some of the factors that cause a chemical reaction to occur. 3. Be able to identify the solid that forms in a precipitation reaction. 4. Be able to write and balance molecular, complete ionic, and net ionic equations for a chemical reaction. 5. Be able to write a balanced chemical equation for the reaction between an acid and a base. 6.

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Balance and classify the following
chemical equation. $\text{Ag}_2\text{SO}_4 + \text{NaCl}$
 $\text{AgCl} + \text{Na}_2\text{SO}_4$... Balancing
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equations, solubility, precipitation,

acid-base, gas evolution,
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Not all single replacement reactions
 actually occur. Reactivity of elements
 determines whether or not the reaction
 takes place. If the reactivity of the single
 element is greater than the reactivity of
 the element in the compound it is
 attempting to replace - the reaction will
 happen.

**9.2 Classifying Chemical Reactions
 Flashcards | Quizlet**

$C(s) + O_2(g) \rightarrow CO_2(g)$ $H_2O(l) + SO_3(g) \rightarrow H_2SO_4(aq)$ Decomposition
 reactions. A single reactant breaks down
 to form 2 or more products. Examples: $H_2CO_3(aq) \rightarrow H_2O(l) + CO_2(g)$ $CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$ Single-
 replacement reactions. A single element
 replaces a similar element of an
 adjacent reactant compound.

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Chapter 5: Classifying and Balancing of Chemical Reactions Part 1: Balancing Chemical Equations Sections: 5.1, 5.2 Goals: 1)

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Chemical reactions occur when one or more molecules change chemically. In other words, bonds break between the reactants, and atoms rearrange to form new chemical bonds. These reconfigured atoms can either be a compound or just an element. The products have different chemical identities compared to the reactants. Common examples of chemical reactions include the burning of fuel, respiration, and the formation of rust.

How to Classify Chemical Reactions for AP® Chemistry ...

Classifying Chemical Reactions:

Chemical reactions can be classified based on the nature by which the reactant molecules are converted to the products. In some cases, reactions can also be ...

Classify the following reaction: **Fe₂(SO₄)₃(aq) - study.com**

Most chemical reactions can be classified into one or more of five basic types: acid-base reactions, exchange reactions, condensation reactions (and the reverse, cleavage reactions), and oxidation-reduction reactions. The general forms of these five kinds of reactions are summarized in Table

7.10.1, along with examples of each.

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Types or Classification of Chemical Reactions - This video goes over the different Types of Chemical Reactions typically taught in college and high school ch...