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Section 10.1 – The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance. 6.02×10^{23} is called Avogadro's number. 1 mole = 6.02×10^{23} representative particles

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Chapter 10 Chemical Quantities. mole. Avogadro's number. representative particle. molar mass. the SI unit for measuring the amount of a substance. number of representative particles in a mole, 6.02×10^{23} . refers to the species present in a substance: usually atoms, m.... the mass of one mole of a

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Chapter 10: Chemical Quantities

CHAPTER 10: Chemical Quantities

BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

Chapter 10 Chemical Quantities

Answer Key

Chapter 10 "Chemical Quantities". the SI unit representing 6.02×10^{23} representative particles of a substance. the temperature and pressure at which

one mole of gas occupies a volume of 22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

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