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each atom: a) SF 2 20 electrons 0 0 0 b) NH 2 OH (N and O are bonded to one another) 14 electrons 0 0 0 0 0 c) PO 4 3-

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Chemistry 101 ANSWER KEY 1 REVIEW QUESTIONS Chapter 10 1. Draw Lewis structures and determine the molecular geometry of each molecule or ion shown below: A) ClO 2 - 20 electrons Bent (from tetrahedral) B) ICl 3 28 electrons T-shaped (from trigonal bipyramidal) C) TeF 4 34 electrons See-saw (from trigonal bipyramidal) 2.

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2p6 3s2 Page 5/26

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General Chemistry 101 Chapter 1 Review Questions 1) Molecules can be described as A) mixtures of two or more pure substances. B) mixtures of two or more elements that has a specific ratio between components. C) two or more atoms chemically joined together. D) heterogeneous mixtures. E)

homogeneous mixtures. Answer: C 2)
Identify the TRUE statement.

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is used to attain specific goals. Always
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of matter and the changes it undergoes.
organic chemistry. study of essentially
all compounds that contain carbon.
inorganic chemistry. study of essentially
all compounds that do not contain
carbon. physical chemistry.

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