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# Network Exam Objectives

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### Covalent Bonding Chapter 8 Worksheet

Chapter 8 Covalent Bonding 183 Section Review Objectives • State a rule that usually tells how many electrons are shared to form a covalent bond • Describe how electron dot formulas are used • Predict when two atoms are likely to be joined by a double or a triple covalent bond • Distinguish between a single covalent bond and other covalent bonds

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Chapter 8 Covalent Bonding and Molecular Structure 8-6 Ex ample: PF 3 Ex CO 2 Step 1: Count Valence Electrons Count the total number of valence electrons in the molecule or ion. Anions have extra electrons, so add 1 electron for each negative charge. Cations have a deficiency of electrons, so subtract 1 electron for each positive charge.

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Grade 8 Ionic Bonding Lewis Dot Structures Some of the worksheets for this concept are Ws lewis structures ionic compounds, Covalent, Chapter 8 covalent bonding and molecular structure, Practice problems h s so ch br hcn, Electron dot lewis structures, University of illinois at urbanachampaign, Chemistry lewis dot structures answers, A guide to chemical bonding.

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Chapter 7 Worksheet Fall 2006 page 4 of 5 Formula ABE formula Number of e-domains on central atom # e- domains/ # non-bonding domains on central atom Electron-Domain Geometry Molecular Geometry (name) Bond angle(s) on central atom CO O 3 CO 3 2-SO 2 PF 5 SF 4 SF 6 XeF 4 11) a. Identify the molecules in the table above that are polar.