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 14. Chemical Kinetics -
 5 - Practice Exercise
 14.3 . The
 decomposition of N_2O_5
 proceeds according
 to the following
 equation: $2\text{N}_2\text{O}_5(\text{g})$
 $4\text{NO}_2(\text{g}) + \text{O}_2(\text{g})$ If
 the rate of
 decomposition of N_2O_5
 at a particular
 instant in a reaction

vessel is 4.2×10^{-7} M/s, what is the rate of appearance of .

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Chapter 14. Chemical Kinetics 14.1 Factors that Affect Reaction Rates • Chemical kinetics = the study of how fast chemical reactions occur. • Four factors which affect rates of reactions: 1. 2. 3. 4. 14.2 Reaction Rates • speed of a reaction = change in [reactant] or [product] per unit time = reaction rate. Average rate = Sample Exercise 14.1 (p. 576)

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14. Chemical Kinetics -
5 - Practice Exercise
14.3 . The
decomposition of N_2O_5
proceeds according
to the following
equation: $2\text{N}_2\text{O}_5(\text{g}) \rightarrow 4\text{NO}_2(\text{g}) + \text{O}_2(\text{g})$ If
the rate of
decomposition of N_2O_5 at a particular
instant in a reaction
vessel is 4.2×10^{-7}
M/s, what is the rate of
appearance of . a) NO_2 .
 $2 (8.4 \times 10^{-7} \dots$

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Chemical Kinetics.
Common Student
Misconceptions. • It is
possible for
mathematics to get in
the way of some
students'.
understanding of the
chemistry of this
chapter. •Students

often assume that reaction orders may be determined from stoichiometric coefficients regardless of the reaction mechanism.

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14.39. a. $\text{HClO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. Only the forward reaction is indicated since HClO_4 is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction is commonly written without water as a reactant. The common abbreviation for this reaction is: $\text{HClO}_4(\text{aq}) \rightarrow \text{H}^+$.

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3 - Instantaneous Rate
• We can plot $[\text{C}_4\text{H}_9\text{Cl}]$ versus time. •
The rate at any instant
in time is called the
instantaneous rate. • It
is the slope of the

straight line tangent to
the curve at that
instant. •
Instantaneous rate is
different from average
rate.

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