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Specific Heat Worksheet 2 Answers

Specific Heat Worksheet #2 Name: ____ Per: ____ Seat: ____
Directions: Calculate the following showing ALL work to receive credit. Formula $Q = mc \Delta T$, where Q is heat in joules, c is specific heat capacity in J/g C, m is the mass in grams, and delta T is the change in temperature in C. Q Work Answer with Units! 1 How much heat is lost when a 640 g piece of copper cools from 375 °C, to 26 °C?

Specific Heat Worksheet #2.pdf - Specific Heat Worksheet#2 ...

Name Answer Key Date 9/9/15 Chp 2-1: Specific Heat Worksheet (m) (ΔT) (C sp)= Q 1. Specific heat is the amount of energy that it takes to raise the temperature of 1 gram of a substance by 1

degree kelvin 2. Absolute zero is the temperature at which all molecular motion ceases 3. Endothermic process is a change in matter in which energy is absorbed 4.

Specific Heat WS Answers - Name Answer Key Date Chp 2-1 ...

Chapter 10 Worksheet #2 1. Calculate the energy require (in calories) to heat 10.4 g of mercury from 37.0 oC to 42.0 oC. Specific heat of mercury is 0.14 J/g oC. $q = m c \Delta t$ $q = 10.4 \text{ g} \cdot 0.14 \text{ J/g oC} \cdot 5.00 \text{ oC} = 7.28 \text{ J} \cdot 1 \text{ cal} = 1.74 \text{ cal}$ 4.184 J 2. If 50. J of heat are applied to 10. g of iron, by how much will the temperature of the iron

Chapter 10 Worksheet #2 Answer

(ANSWERS) 1. A 500 g piece of iron changes 7°C when heat is added. How much heat energy produced this change in temperature? (Ans. 2,000 J) 2. When 300. cal of energy is lost from a 125 g object, the temperature decreases from 45.0°C to 40.0°C. What is the specific heat of this object? (Ans. 0.48 cal/g oC or 2.0 J/g oC)

Honors Chemistry Worksheet - Specific Heat

Before discussing Calculating Specific Heat Worksheet Answers, you need to recognize that Knowledge can be your answer to a better the next day, along with studying doesn't just stop the moment the school bell rings. Of which getting claimed, many of us provide you with a a number of basic yet helpful posts along with design templates made ideal for almost any educative purpose.

Calculating Specific Heat Worksheet Answers | akademiexcel.com

q = amount of heat (J) m = mass (grams) c = specific heat (J/g°C) ΔT = change in temperature (°C) 2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other. Heat is a combination of kinetic energy (measured by temperature) and potential energy. a. Perform calculations using: ($q = m c \Delta T$) b.

Worksheet- Calculations involving Specific Heat

Answers are provided at the end of the worksheet without units. 1. A 15.75-g piece of iron sorbs 1086.75 joules of heat energy, and its temperature changes from 25 0 1750C. Calculate the specific heat capacity of iron. = 'C ' Q 5) 2. How many joules of heat are neeaea Fo ratse the temperature of 10.0 g of

Specific Heat Wksht20130116145212867

Two page worksheet using Specific Heat Capacity. Questions start easy then become gradually harder. Answers included on separate sheet. Also includes a spreadsheet to show how the calculations have been done.

Specific Heat Capacity Worksheet (with answers) | Teaching ...

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Specific Heat Problems Worksheet Answers Along with Specific Heat Worksheet Answers. You can get a sheet that will help you with all the basic needs for an air conditioning system. When you are looking for a sheet, you can also check online to determine what works best for you.

Specific Heat Problems Worksheet Answers

Heat Transfer/ Specific Heat Problems Worksheet Solving For Heat (q) 1. How many joules of heat are required to raise the temperature of 550 g of water from 12.0 °C to 18.0 °C? 2. How much heat is lost when a 64 g piece of copper cools from 375 °C, to 26 °C? (The specific heat of copper is 0.38452 J/g °C). Place your answer in kJ. 3.

Heat Transfer/ Specific Heat Problems Worksheet

For the last step, with proper sig figs, I get 91.2, which is essentially the same answer as step #3. This is because of the small specific heat of the aluminum (0.089 J/g °C). 19. A sample of cobalt, A, with a mass of 5.00 g, is initially at 25.0 °C. When this sample gains 6.70 J of heat, the temperature rises to 27.9 °C.

Thermochem Worksheet #2 Answers - ChemTeam

What is the specific heat of an unknown substance if a 2.50 g sample releases 12 calories as its temperature changes from 25°C to 20°C? ANSWER KEY. HEAT Practice Problems . $Q = m \times$

$\Delta T \times C$. 5.0 g of copper was heated from 20°C to 80°C. How much energy was used to heat Cu? (Specific heat capacity of Cu is 0.092 cal/g °C) 27.6 cal

HEAT Practice Problems

Worksheet- Calculations involving Specific Heat 1. For $q = m \times c \times \Delta T$: identify each variables by name & the units associated with it. 2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other. (-m.c.AT) a. Perform calculations using 1. Gold has a specific heat of 0.129 J/(g°C). How

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Specific Heat and Heat Capacity Worksheet DIRECTIONS: Use $q = (m)(C_p)(\Delta T)$ to solve the following problems. Show all work and units. Ex: How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C? 1.

Specific Heat and Heat Capacity Worksheet

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Campbell Science - Home

For $q = m \times c \times \Delta T$: identify each variables by name & the units associated with it. q = amount of heat (J) m = mass (grams) c = specific heat (J/g°C) ΔT = change in temperature (°C) 2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other.

Chemistry Specific Heat Worksheet Answers

Here are the heat capacities of the four substances: 0.10 cal/g °C, 0.25 cal/g °C, 1.0 cal/g °C, & 0.2 cal/g °C. Match & then label each substance with its specific heat capacity on the graph. See graph

above. 7. If something has a high specific heat capacity will it take a lot of heat or a little heat to change its temperature? Explain ...