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SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505–510)

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$\Delta H^\circ = \Delta E^\circ + (Dn)RT$ $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(l)$ $0.008314 \text{ kJ H} = 222 \text{ kJ} + (0.1.5) \text{ mol } 298.15 \text{ K} = \text{K mol. } \Delta^\circ - - - 226 \text{ kJ/mol } H_2$. 3. The heat of combustion of liquid cyclohexane, $C_6H_{12}(l)$, is -3924 kJ/mole . 8.25 g of cyclohexane is. placed in the bomb of a bomb calorimeter with excess oxygen.

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