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Chapter 11 Chemical Reactions Page Section 11.1 - Describing Chemical Reactions In a chemical reaction, the reactants are written on the left and the products on the right. The arrow that separates them is called yield. Reactants $\rightarrow$ Products

Chapter 11: Chemical Reactions

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Write a balanced chemical equation for each reaction. Use the necessary

symbols from Table 11.1 to describe the reaction completely. a. Bubbling chlorine gas through a solution of potassium iodide gives elemental iodine and a solution of potassium chloride. b. Bubbles of hydrogen gas and aqueous iron (III) chloride are produced when metallic

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Chapter 11 - Chemical Reactions - 11 Assessment - Page 380: 67

SECTION 11.2 TYPES OF CHEMICAL REACTIONS (pages 330-339) This section explains how to identify a reaction as a combination, decomposition, single-replacement, double-replacement, or combustion reaction. It also describes how to predict the products of each type of reaction. Classifying Reactions (page 330)

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321-329)

Chapter 11: Reactions and Other

Chemical Processes Expand/collapse
global location Chapter 11 Problems Last
updated; Save as PDF Page ID ... Assume
there are no side reactions or auxiliary
reactions. From Eqs. 11.5.9 and 11.5.10,
calculate the standard molar internal
energy of combustion of n-hexane at
(298.15 K). (p) ...

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11 Chemical Reactions Chemical
Equation n Describes chemical reaction.
n Chemical

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Name ____ Date ____ Chapter 11 Test:
Stoichiometry 1. Write a balanced
chemical equation for a reaction
between zinc and copper II sulfate. 2. If
5.00 grams of zinc reacts with 5.00
grams of copper II sulfate, determine the
limiting reactant. 3.

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Chapter 11 Chemical Reactions
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Types of Chemical Reactions> 13 A
decomposition reaction is a chemical
change in which a single compound
breaks down into two or more simpler
products. • Decomposition reactions

involve only one reactant and two or

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Figure 11.3.1 An Ammonium Dichromate
Volcano: Change during a Chemical
Reaction The starting material (left) is
solid ammonium dichromate. A chemical
reaction (right) transforms it to solid
chromium (III) oxide, depicted showing a
portion of its chained structure, nitrogen
gas, and water vapor.

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Write correct formulas for the products
in these double replacement reactions.
1) $\text{H}_3\text{PO}_4 + 2\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow \text{Ba}_3(\text{PO}_4)_2 + 6\text{KCl}$
2) $\text{AgNO}_3 + \text{KCl} \rightarrow \text{AgCl} + \text{KNO}_3$
3) $\text{H}_2\text{SO}_4 + \text{Co} \rightarrow \text{CoSO}_4 + \text{H}_2$
4) $\text{BCl}_3 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{BO}_3 + \text{HCl}$
5) $\text{Ag}_2\text{C}_2\text{H}_3\text{O}_2 + \text{K}_2\text{CrO}_4 \rightarrow \text{K}_2\text{C}_2\text{H}_3\text{O}_2 + \text{Ag}_2\text{CrO}_4$

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Section 11.2 Types of Chemical
Reactions 331 CONCEPTUAL PROBLEM
11.4 Writing Equations for Combination
Reactions Copper and sulfur, shown in
the photo, are the reactants in a
combination reaction. Complete the
equation for the reaction.