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Kinetics, Dynamics and Mechanism - Department of Chemistry

The solution is actually very simple: the reaction rate is defined as the rate of change of the concentration of a reactant or product divided by its stoichiometric coefficient. For the above reaction, the rate (usually given the symbol v) is therefore.

$$v = - \frac{1}{2} \frac{d[N_2]}{dt} = \frac{1}{3} \frac{d[H_2]}{dt} = \frac{1}{2} \frac{d[NH_3]}{dt}$$

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