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# Return To Sender Book

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### Chemquest 24 Answer Key

3. Another example:  $7.45 \times 10^{-8} = 0.000000745$  the “-8” tells you that there are 8 decimal places between the right side of the 7 and the end of the number.

### ChemQuestsanswers - Yumpu

24 ChemQuest 8 Name: \_\_\_\_ Date: \_\_\_\_ Hour: \_\_\_\_ Information: Structure of the Atom Note the following symbols: (they are not to scale) = proton (positive charge) = electron (negative charge) = neutron (no charge) The following three diagrams are hydrogen atoms: 1 1H 2 1H 3 1H The following three diagrams are carbon atoms: Notice the type of notation used for atoms: 12 13 14 6C 6 C 6C (6 ...

### 24 ChemQuest 8 Name:

Hint: add your answer to question 1 and four times your answer to question 2.  $n \times t \times (q \times BZ \times e \times \text{AronS})$  Verify that there are 32 electrons pictured in the drawing of  $\text{CCl}_4$ . 3es Find the sum of all the valence electrons for  $\text{NF}_3$ . (Add how many valence electrons one nitrogen atom has with the valence electrons for three fluorine atoms.)  $\text{H} \times \text{crl} \times 2 \times \text{Tte} \times \text{Je}$ .

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PLEASE NOTE: If you have a question about these answers, it is your responsibility to come to office hours or ask during class work time. UNIT 12 - HW Practice Keys - ChemActivity 51: Cell Voltage - ChemActivity 50: Electrochemical Cell UNIT 11 - HW Practice Keys - ChemQuest 55: Free Energy - ChemQuest 54: 2nd Law of Thermodynamics

**HW Keys - Roosevelt High School AP Chemistry 2017-18**

Answer: + that S has a negative 2 charge. There are two "X" ions to balance out the negative 2 charge of S. Step 2: Criss cross to balance charges. Step 1: Write each ion and each charge. Positive first! Here's what you get: Add parentheses balance charges. Step 2. Criss cross to

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Hint: add your answer to question 1 and four times your answer to question 2.  $4 + 4(7) = 4 + 28 = 32$  4. Verify that there are 32 electrons pictured in the drawing of  $\text{CCl}_4$ . Count each dot in the picture above: there are 24. There are also 8 more electrons represented by the 4 lines (each line represents two electrons). 5.

**ChemQuest 23 : Drawing Covalent Compounds**

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