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**Worksheet On Balancing Chemical
Equations**

Both of them are separated by an arrow.

For instance, $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ denotes that there are four atoms of hydrogen and 2 atoms of oxygen on both sides of the equation. The amount of reactants must be equal to the amount of products. When students get big chemical equations in a balancing equation worksheet, they often find it to be very difficult. We will help you understand through some tips in this article too, to help you get through the process seamlessly.

49 Balancing Chemical Equations Worksheets [with Answers]

An Easy Example To Get You Started
With Step 1: Start by counting the number of atoms present, for each element on the side of the reactants as well as the... Step 2: When you find out

that certain elements are not balanced, place the required coefficient that is needed to... Step 3: Once you are ...

100 Balancing Chemical Equations Worksheets with Answers ...

Balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow. This is a collection of printable worksheets to practice balancing equations. The printable worksheets are provided in pdf format with separate answer keys.

How to Balance Equations - Printable Worksheets

Balancing Chemical Equations
Worksheet 1 $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ 2. Na

$+ \text{Cl}_2 \rightarrow \text{NaCl}$ 3. $\text{N}_2 + \text{O}_2 \rightarrow \text{NO}_2$ 4. $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$ 5. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ 6.
 $\text{Ca} + \text{N}_2 \rightarrow \text{Ca}_3\text{N}_2$ 7. $\text{Li} + \text{F}_2 \rightarrow \text{LiF}$ 8.
 $\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$ 9. $\text{NH}_3 \rightarrow \text{N}_2 + \text{H}_2$ 10. $\text{HCl} \rightarrow \text{H}_2 + \text{Cl}_2$ 11. $\text{NI}_3 \rightarrow \text{N}_2 + \text{I}_2$ 12. $\text{HI} \rightarrow \text{H}_2 + \text{I}_2$ 13. $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 14. $\text{K} + \text{S} \rightarrow \text{K}_2\text{S}$ 15. $\text{Li} + \text{Cl}_2 \rightarrow \text{LiCl}$

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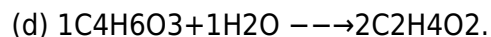
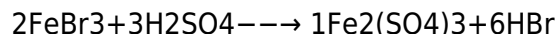
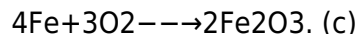
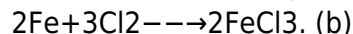
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(Coefficients equal to one (1) do not need to be shown in your answers). (a)



Balancing Equations: Practice Problems

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