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 $(90 - 20) = 14000 \text{ cal}$. $Q_{\text{water}} = 1000.1$.
 $(90 - 20) = 70000 \text{ cal}$. $Q_{\text{calorimeter}}$
 $= 70000 + 14000 = 84000 \text{ cal}$. 1 mol S is
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Thermochemistry Thermochemistry and Energy and Temperature
Thermochemistry is study of changes in energy (heat) associated with physical or chemical changes. Force = push $F = m a$ (mass x acceleration) force units: N (newton) = kg m s⁻² Work = force x distance $W = F d$ energy units: J (joule) = kg m² s⁻²

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$\Delta H = M \times C \times -\Delta T$. Where M is the mass

or volume of the solution being used to test the energy change, C is the Specific Heat Capacity of the solution, (for example the specific heat capacity of water is 4.1855 J g⁻¹ K⁻¹), and ΔT is the change in temperature.

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Thermochemistry is a branch of Thermodynamics, a scientific field that comprises the relationships between heat, work, and other forms of energy resulting from different chemical and physical processes. Thermochemistry itself is defined as energy changes when a chemical reaction takes place. Energy is the capacity to supply heat or do work.

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The answer is given per mole of reaction. $\Delta_f H^\circ$ (298 K) for SF_4 (g), H_2

O (l), SO_2 (g) and HF (g) are -763 , -286 , -297 and -273 kJ mol^{-1} , respectively. What is the value of $\Delta_r H^\circ$ (298 K) for the hydrolysis shown above? For this process, $\Delta_c H^\circ$ (298 K) = -3953 kJ mol^{-1} .

Chapter 2: Thermochemistry

42. Given that heat neutralization of strong acid and strong base is -13.7 kcal. The heat produced when one mole of HCl is mixed with 0.5 mole of NaOH will be _____